**Chemistry I**

**Grading:**

A (93% - 100%) = 4 grade points

B (85% - 92%) = 3 grade points

C (75% - 84%) = 2 grade points

D (70% - 74%) = 1 grade point

F (0% - 69%) = 0 grade points

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| * **End-Of-Course (EOC) exam in Chemistry** | 25% of the 2nd semester grade; the EOC is a state test given in May and cannot be retaken for a better score or made up if absent; the EOC is cumulative (includes all curriculum standards for the year); classwork, homework, and quizzes contain practice EOC questions |
| * **semester exam** | tests the cumulative knowledge for the semester; contains practice for the EOC; cannot be retaken for a better grade |
| * **quizzes** | contain approximately a week’s-worth of classwork and practice questions for the EOC; can be made up if absent; can be re-taken for a better grade (average of the new and old grades) only if failed |
| * **projects** | contain practice for the EOC; can be made up if absent |
| * **classwork** | contains practice for the EOC; classwork collected daily; can be made up if absent |
| * **homework** | contains practice for the EOC; homework can be made up if absent |

**Contacting Ms. Kirby:**

* in person at school before or after school
* by phone at (901) 416-4540
* by e-mail at [KIRBYJ@scsk12.org](mailto:KIRBYJ@scsk12.org)
* website www.kirbyscience@ yola.com

**Attendance:**

* Mandatory 100% attendance policy.
* Each student is required to come on time and be prepared for class (**textbook,** **paper & 2 writing instruments**).
* Each student is required to remain in class for the entire class period and to take care of restroom or water needs before class.

**Disciplinary Class Rules:**

All students have the right to learn and the right to safety. All students have the responsibility to respect others and themselves.

1. Be in your seat when the bell rings
2. Be prepared for class (textbook, paper, 2 ready writing instruments, water & restroom needs already taken care of)
3. Respect your environment (no littering or graffiti)
4. Respect others (no talking to disrupt learning; keep your opinion of others to yourself = no checking or bullying or profanity)
5. Respect school rules (cell phones/electronics off & out of sight; no food/drink - including water bottles; no makeup or lotion; in dress code)

**Makeup Policy:**

* Students must request and complete make-up for missed work and quizzes or receive 0 points for all missed work.
* Standard (not Honors!) students can retake a weekly quiz for a better grade (average of the new and old test grade) only if failed.

**Supplies:**

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| **Bring-To-Class** Everyday | **Leave-In-Class** With Teacher For Use During School Year |
| * pencil or pen (sharpened & in working order) | * notebook paper (1 packet) |
| * folder or binder with dividers (syllabus, notes, and graded work) | * black dry erase markers (2 markers) |
| * notebook paper * calculator (Chemistry only) * 1-subject spiral notebook (stays in class) | * safety blunt scissors (1) * white school glue (1 bottle) * facial tissue (2 boxes) * unscented baby wipes (1 refill) |

**Classroom Procedures:**

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| **ACTIVITY** | **CONVERSATION**  **between students** | **HELP**  **from teacher** | **IN-CLASS MOVEMENT**  **(sharpening pencils, getting tissue)** | **PARTICIPATION** |
| **teacher lecture** | **none** | **raise hand** | **none** | **listen & take notes only** |
| **group project or individual class work** | low-voice only;  in-group or shoulder-partners;  on-task only | 1st in-group or shoulder-partner;  2nd raise hand | open (ask teacher for permission) | on-task only |
| **lab** | low-voice only; in-group only;  on-task only | 1st in-group;  2nd raise hand | restricted (as instructed only) | in-group only;  on-task only |
| **quiz or exam** | **none** | **none** | **none** | **testing only** |

**Curriculum Standards for Chemistry I:**

***First Quarter***

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| **STATE STANDARDS** | **EMBEDDED STANDARDS** |
| CLE 3221.2.1 Investigate the characteristic properties of matter. | 3221.Inq.2 Identify an answerable question and formulate a hypothesis to guide a scientific investigation. |
| CLE 3221.2.2 Explore the interactions between matter and energy. | 3221.Inq.4 Perform and understand laboratory procedures directed at testing hypothesis. |
| CLE 3221.2.3 Apply the kinetic molecular theory to describe solids, liquids, and gases. | CLE 3210.Math.2 Utilize appropriate mathematical equations and processes to understand chemistry concepts. |
| CLE 3221.1.1 Compare and contrast historical models of the atom. | |  | | --- | | CLE 3210.Math 1 – Understand the mathematical principles associated with the science of chemistry. | |
| CLE 3221.1.3 Describe an atom in terms of its composition and electron characteristics. | 3221.Inq.5 Select appropriate tools and technology to collect precise and accurate quantitative and qualitative data. |
|  | 3221.Inq.6 Correctly read a thermometer, balance, metric ruler and a graduated cylinder. |
| 3221.Math.1 Use a variety of appropriate notations (e.g., exponential, functional, square root). |
| 3221.Math.5 Analyze graphs to describe the behavior of functions. |
| 3221. Inq.7 Record observation and/or data using correct scientific units and significant figures. |
| 3221.Inq.1 Trace the historical development of a scientific principle or theory. |

***Second Quarter***

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| **STATE STANDARDS** | **EMBEDDED STANDARDS** |
| CLE 3221.1.2 Analyze the organization of the modern periodic table. | CLE 3210.Inq.5 Compare experimental evidence and conclusions with those drawn by others about the same testable question. |
| CLE 3221.3.1 Investigate chemical bonding. | CLE 3210.T/E.2 Differentiate among elements of the engineering design cycle: design constraints, model building, testing, evaluating, modifying, and retesting. |
| CLE 3221.3.2 Analyze chemical and nuclear reactions | CLE 3210.Inq.1 Recognize that science is a progressive endeavor that reevaluates and extends what is already accepted. |
| CLE 3221.3.3 Explore the mathematics of chemical formulas and equations. | CLE 3210.Inq.3 Use appropriate tools and technology to collect precise and accurate data. |
| CLE 3221.3.4 Explain the law of conservation of mass/energy. | CLE 3221.Math.1 Understand the mathematical principles associated with the science of chemistry. |

***Third Quarter***

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| **STATE STANDARDS** | **EMBEDDED STANDARDS** |
| CLE 3221.3.3 Explore the mathematics of chemical formulas and equations. | CLE 3221.Math.1 Understand the mathematical principles associated with the science of chemistry. |
| CLE 3221.3.4 Explain the law of conservation of mass/energy | CLE 3210.Inq.6 Communicate and defend scientific findings. |
| CLE 3221.2.4 Investigate characteristics associated with the gaseous state. | CLE 3221.Math.2 Utilize appropriate mathematical equations and processes to solve chemistry problems. |
| CLE 3221.2.3 Apply the kinetic molecular theory to describe solids, liquids, and gases. | CLE 3210.Inq.4 Apply qualitative and quantitative measures to analyze data and draw conclusions that are free of bias. |
| CLE 3221.2.5 Discuss phase diagrams of one component systems. |  |

***Fourth Quarter***

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| **STATE STANDARDS** | **EMBEDDED STANDARDS** |
| CLE 3221.2.2 Explore the interactions between matter and energy. | CLE 3210.T/E.4 Describe the dynamic interplay among science, technology, and engineering within living, earth-space, and physical systems. |
| CLE 3221.3.4 Explain the law of conservation of mass/energy | CLE 3210.T/E.1 Explore the impact of technology on social, political, and economic systems. |
| CLE 3221.3.2 Analyze chemical and nuclear reactions | CLE 3221.Math.1 Understand the mathematical principles associated with the science of chemistry. |
|  | CLE 3221.Math.2 Utilize appropriate mathematical equations and processes to solve chemistry problems. |